

Name \_\_\_\_\_

AP Chemistry

## Molecular Geometry & Polarity Practice

- 1) For each of the following compounds, draw the Lewis structures, predict the molecular geometry, and predict if the molecule is polar on a separate sheet of paper.
- (a)  $\text{OCl}_2$
  - (b)  $\text{KrF}_2$
  - (c)  $\text{BeH}_2$
  - (d)  $\text{SO}_2$
  - (e)  $\text{SO}_3$
  - (f)  $\text{NF}_3$
  - (g)  $\text{IF}_3$
  - (h)  $\text{IF}_5$
  - (i)  $\text{SeF}_4$
  - (j)  $\text{CF}_4$
  - (k)  $\text{KrF}_4$
  - (l)  $\text{AsF}_5$
  - (m)  $\text{H}_2\text{O}$
  - (n)  $\text{CO}_2$
  - (o)  $\text{XeF}_2$
  - (p)  $\text{BF}_3$
  - (q)  $\text{ClF}_3$
  - (r)  $\text{CH}_4$
  - (s)  $\text{SF}_4$
  - (t)  $\text{XeF}_4$
  - (u)  $\text{BrF}_5$
  - (v)  $\text{PCl}_5$
  - (w)  $\text{SF}_6$